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ALGORITHMS AND THEIR APPLICATION IN CHEMISTRY

Abstract. *The article examines how chemical processes and problems can be modeled algorithmically using algorithms, which are applied to applied mathematics, in particular to chemical processes, and their significance is revealed.*

Keywords: *algorithm, sequence, chemistry, chemical processes, iterative, branching, linear, modular algorithms modeling, number of moles, molar mass, temperature.*

Algorithm By "routine" is usually meant a set of precise, sequential, and repeatable operations. Algorithms in higher mathematics are essentially structured this way. They are a set of steps designed to solve a specific mathematical problem, each step of which is clearly defined.

We will look at the types of algorithms and their applications to chemistry.

1. Linear algorithm (sequential algorithm)

► □ Definition: In the algorithm, all operations are performed sequentially, there is no branching or repetition.

✓ Features:

- Each step is performed in a specific order.
- There are no conditions.



Application of linear algorithms in chemistry.

Issue: Calculate how many grams of a substance is equal to 1 mole.

Formula: $\text{mass} = \text{number of moles} \times \text{molar mass}$

This calculation is performed sequentially, i.e. it is an example of a linear algorithm.

Steps of the algorithm (in linear form):

- 1) Enter molar mass (M)
- 2) Enter the number of moles (n)
- 3) Mass calculation: $m = n \times M$
- 4) Output

Example: How many grams are 2 moles of water (H_2O)?

Solution:

- 1) Molar mass (H_2O)

$$\text{H} = 1 \times 2 = 2 \text{ g/mol}$$

$$\text{O} = 16 \text{ g/mol}$$

$$\text{General: } M = 18 \text{ g/mol}$$

- 2) Number of moles

$$n = 2 \text{ moles}$$

- 3) Calculation

$$m = n \times M = 2 \times 18 = 36 \text{ grams}$$

Conclusion: In this calculation, all operations were performed sequentially, there were no conditions or repetitions — so this is a linear algorithm.

2. Iterative algorithm (iterative algorithm)

►□Definition: In an algorithm, an action or group of actions is repeated a number of times. The number of repetitions is predetermined or determined by a condition.

✓Features:

- There will be loops in the form of "For", "While" or "Repeat...Until".



Application of iterative algorithms in chemistry.

Problem: During the experiment, the temperature increases by 2°C every 5 minutes. If the initial temperature is 20°C, calculate how the temperature changes over 1 hour.

In this problem, the temperature increases every 5 minutes, meaning that one action is repeated continuously — this is a recursive algorithm.

Algorithm steps:

- 1) Setting the initial temperature:

$T = 20^{\circ}\text{C}$

- 2) Starting time at 0:

$t = 0$

- 3) Recurring action (1 hour = 60 minutes, every 5 minutes):

- Every time $T = T + 2$
- $t = t + 5$

- 4) Output the result at each stage (optional)

- 5) Repeat: until the time reaches 60 minutes

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0	20
5	22
10	24
15	26
...	...
60	44

1 soat oxirida harorat: 44°C

Conclusion: In this problem, the same action (increasing the temperature by 2°C) is repeated every 5 minutes — this is an example of a classic iterative algorithm.

3. Branching algorithm (conditional algorithm)

► □ **Definition:** In an algorithm, a path is chosen based on a condition (if-else).

✓ **Features:**

- The decision is made.
- Different paths are chosen in different situations.



Application of branching (conditional) algorithm in chemistry.

Issue: Determining the physical state of matter.

Condition: The temperature of the water is given. Depending on it:

- If the temperature is $< 0^{\circ}\text{C}$ — ice (solid)
- If $0^{\circ}\text{C} \leq \text{temperature} < 100^{\circ}\text{C}$ — liquid
- If the temperature is $\geq 100^{\circ}\text{C}$ — steam (gas)

Here, there are conditions and different paths are chosen based on them — this is a clear example of a branching algorithm.

Algorithm steps:

- 1) Enter the temperature value (T)
- 2) Check the condition:
 - If $T < 0 \rightarrow$ "Physical state: Ice (solid)"
 - Otherwise if $T < 100 \rightarrow$ "Physical state: Liquid"
 - Otherwise $\rightarrow$ "Physical state: Vapor (gas)"

Misollar bilan:

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-5	Muz (qattiq)
25	Suyuqlik
100	Bug' (gaz)

Conclusion: In this case, conditional decisions are being made depending on the temperature. The result changes depending on the condition. So this is a branching (conditional) algorithm.

4. Recursive algorithm (self-calling)

► □ **Definition:** The algorithm calls itself to execute a smaller version of itself. Each call brings the result closer to the final state.

✓ **Features:**

- Each time it calls itself with a smaller subproblem.
- Once it reaches the base position, it begins to return.



Application of recursive algorithm in chemistry.

Issue: To understand the levels of chemical elements, calculate the number of electron shells from the number of protons (in a simple model).

A simple rule-based model (for understanding):

Let's assume that the electrons are arranged in layers as follows:

- 1st floor: 2 electrons

- 2nd layer: 8 electrons
- Layer 3: 8 electrons
- and so on...

Given the number of electrons, we recursively determine how many shells are filled.

Example: If there are 13 electrons - how many shells are filled?

For example: $e = 13$

- $13 > 10$, so $1 + \text{floor_number}(13 - 8 = 5)$
- $5 \leq 10 \rightarrow$ goes to the 2nd floor
- Result: $1 + 2 = 3$ floors

Conclusion: Here, the `layer_number` function is repeatedly calling itself, decreasing the number of electrons. This is the simplest and most chemistry-friendly version of the recursive algorithm.

5. Modular algorithm

►□ The main principle of a modular algorithm is that each module performs a separate task.



Application of modular algorithm in chemistry.

Issue: Dividing the chemical calculation process into small modules.

Example: Calculating the mass of a chemical substance.

Problem: To find the mass of a substance, we know the molar mass and the number of moles of the substance.

Modules:

1) **Molar mass calculation module**

(Finds the total molar mass by entering the atomic mass of the elements and their number)

2) **Mass calculation module**

(Mass is calculated by giving molar mass and number of moles)

3) **Output module**

Example: List of elements:

- H: 2, atomic mass 1 g/mol
- O: 1, atomic mass 16 g/mol

Number of moles: 3 moles

Calculation:

- Molar mass = $(2 \times 1) + (1 \times 16) = 18$ g/mol
- Mass = $18 \times 3 = 54$ grams.

Conclusion: Here, each module performs a separate task — this was the basic principle of the modular algorithm. In chemistry, dividing complex calculations into smaller parts and modules makes calculations easier.

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